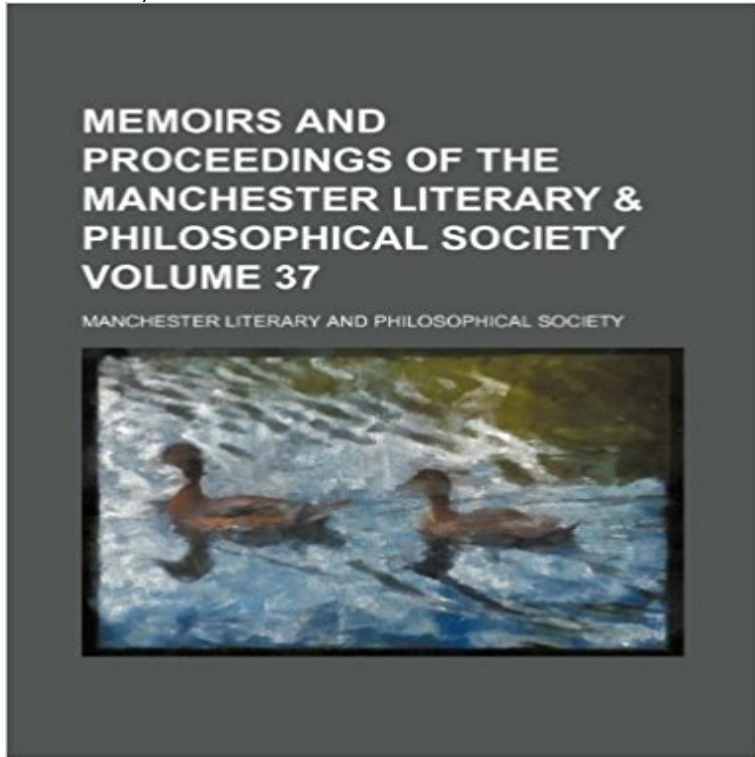


Memoirs and proceedings of the Manchester Literary & Philosophical Society Volume 37



This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1893 Excerpt: ...chemists has opened a new era in the theory of explosions. In July, 1881, two papers appeared in the Comptes Rendus, one by M. Berthelot, the other by MM. Mallard and Le Chatelier. Both papers announced the discovery of the enormous velocity of explosion of gaseous mixtures. Other papers by the same authors quickly followed. M. Berthelot made the important discovery that the rate of explosion rapidly increased from the point of origin until it reached a maximum which remained constant, however long the column of gases might be. This maximum M. Berthelot stated to be independent of the pressure of the Phil, Trans., 1884, Pt. II., p. 635. gases, of the material of the tube, and of its diameter above a small limit. The rate of explosion thus forms a new physico-chemical constant, having important theoretical and practical bearings. The name londe explosive was given by Berthelot to the flame when propagated through an explosive mixture of gases at the maximum velocity. While Berthelot, associated with Vieille, was measuring the rate of the explosion-wave for various mixtures of gases, Mallard and Le Chatelier continued the study of the preliminary phenomena of explosion which precede the formation of the wave. They showed by photographing on a revolving cylinder:--(i) that when a mixture such as nitric oxide and carbon bisulphide is ignited at the open end of a tube, the flame travels a certain distance (depending on the diameter and length of the tube) at a uniform velocity; (2) that at a certain point in the tube, vibrations are set up which alter the character of the flame, and that these vibrations become more intense, the flame swinging backwards and forwards,

with oscillations of increasing amplitude;
and (3) that th...

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